

## Production of Barium Titanate Compounds

### Field of the Invention

The invention relates to the production of ultrafine powders of barium titanate  $\text{BaTiO}_3$  as well as solid solutions and doped compounds derived from barium titanate.

### Background of the Invention

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Barium titanate is an important material for the electronic industry due to its dielectric and piezoelectric properties which make it suitable for many applications such as multilayer ceramic capacitors, positive temperature coefficient (PTC) thermistors, piezoelectric actuators, passive memory storage devices, acoustic transducers, and in  
10 electroluminescent panels.

The market for multilayer ceramic capacitors (MLCC) continues to demand smaller size components for a given capacitance or increased capacitance within a given part design. The capacitance per unit volume for an MLCC can be derived as:

$$C_v = n K_0 K / t^2$$

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where  $C_v$  is the capacitance per unit volume of the active part of the capacitor,  $n$  is the number of active layers,  $K_0$  is the permittivity of free space (a constant),  $K$  is the dielectric constant of the ceramic and  $t$  is the active thickness separating the active layers. In a given MLCC design with a fixed active volume  $C_v$  has to be increased by:

- (1) Using a ceramic with higher  $K$
- 20 (2) Making thinner layers,  $t$
- (3) Using more layers,  $n$ , in the component
- (4) A combination of the above

Option (2) is an effective way to increase  $C_v$  and thinner layers of both ceramic and electrode are required to realize the benefit of more layers (3).

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For use in smaller size MLCC as described above, the following characteristics are required for barium titanate powder:

- Small particle size
- Isotropic shape
- Narrow particle size distribution with limited agglomeration
- 30 - No agglomerates bigger than the thickness of the dielectric layers of the MLCC.
- Controlled stoichiometry with ideal Ba/Ti ratio very close to 1
- High purity
- High density tetragonal phase

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Small particle size is required to design MLCC with dielectric layers smaller than 1 micron in thickness. At least 5 pieces of grains are required between the electrodes to ensure long term stability to the final product; therefore fine grains of less than 200nm are required for a ceramic layer, which is as thin as 1 micron or less. Small

particle size also allows sintering at lower temperature than the usual range 1250-1350°C (this behaviour is also useful when the ceramic is co-fired with palladium-silver alloys or nickel or copper which have a low melting temperature).

5 A narrow particle size distribution allows achieving porosity free and homogeneous microstructure after sintering the ceramic layer. It is well known that narrow particle size distributions and well-dispersed suspensions are required to manufacture pastes for screen-printing or tape casting. Agglomerates bigger than the thickness of the dielectric layers of the MLCC are detrimental to obtaining adequate MLCC.

10 Controlled stoichiometry with Ba/Ti ratio close to 1.000 is necessary to obtain excellent electrical properties and well-defined sintering behaviour. The Ba/Ti ratio also affects the sintering temperature. The presence of excess  $\text{TiO}_2$ , as it forms a eutectic liquid phase with  $\text{BaTiO}_3$  at 1332°C, is known to dramatically influence the microstructure development. The presence of secondary phase  $\text{Ba}_6\text{Ti}_{17}\text{O}_{40}$  leads to the degradation of the insulation resistance at high temperature. Decreasing Ba/Ti ratio  
15 from 0.999 to 0.993 can lead to grain sizes twice bigger after sintering.

High purity powder is required to control the dielectric properties of barium titanate.

20 Better dielectric properties are obtained from high-density barium titanate with a tetragonal structure. Tetragonal barium titanate powders have a better sintering behaviour, higher intrinsic density and lead to ceramics with higher dielectric constants than cubic-phase powders. Also, cubic-phase barium titanate produced by the hydrothermal route exhibits a bloating effect when the MLCC is sintered, causing severe cracks and delamination, as a result of the segregation of porosity in the electrode region.

25 Various methods are used for the production of barium titanate. In order to obtain above mentioned required characteristics, an economically acceptable process or method of production of barium titanate should use low cost equipment and raw material, limit grinding or pulverizing processes, use low temperature, and have short production time.

30 Traditionally barium titanate is prepared using the solid-state route by ball milling  $\text{BaCO}_3$  and  $\text{TiO}_2$ . The mixture is then heated at a temperature of 900°C or above. However,  $\text{BaTiO}_3$  prepared at high temperature exhibits some drawbacks, such as large particle size, a wide size distribution, and a high impurity content resulting from repetitive calcinations and grinding treatments. Further limitations include the inability to control stoichiometry and crystallinity. Often an undesired secondary  
35 phase such as  $\text{BaTi}_2\text{O}_5$  is produced.

Other techniques, such as precipitation, gas condensation, and sol-gel processing, have been developed to improve size, size distribution, and purity. As an example, the sol-gel processing using metal alkoxides as reactants leads to fine particles but

alkoxides are generally too expensive. The oxalate co-precipitation route consisting of precipitation of a mixed barium-titanium oxalate precursor by reaction between a  $\text{BaCl}_2$  and  $\text{TiCl}_4$  solution and an oxalic acid solution requires heat-treatment ( $\sim 700^\circ\text{C}$ ) with the common drawbacks of large particle size, wide particle size distribution and agglomeration.

Hydrothermal synthesis is another wet chemical route to prepare barium titanate. Usually a barium soluble salt (e.g. chloride, nitrate, acetate) or barium hydroxide is used as barium precursor. Titanium is usually provided by titanium dioxide or amorphous titanium hydroxide (prepared by hydrolysing organic or inorganic titanium compounds). Often,  $\text{NaOH}$  or  $\text{KOH}$  are used as mineralising agent. The reactants are subjected to high temperatures and pressures in an autoclave. Using amorphous titanium hydroxide generally leads to a wider particle size distribution and agglomerates, and requires higher cost raw materials. Using  $\text{TiO}_2$  as precursor leads to a narrower particle size distribution but it is difficult to achieve an adequate  $\text{Ba/Ti}$  ratio and the necessary longer production time increases the level of agglomeration and the cost of the process. In general in a large hydrothermal reactor it is difficult to ensure homogeneous nucleation conditions within the whole vessel, which in turn induces a wide size distribution.

Another drawback of a hydrothermal synthesis at moderate temperatures is that it leads to cubic-phase barium titanate with a low density. Hydrothermal conditions for producing commercial quantities are usually under maximum temperature of  $350^\circ\text{C}$  and pressures below 100 Mpa, above which conditions are considered severe (Riman et al., Hydrothermal crystallization of ceramics, *Ann. Chim. Sci. Mat.*, 27(6), pp 15-36, 2002). However, in order to obtain a higher density tetragonal phase barium titanate using the hydrothermal process, a temperature of above  $400^\circ\text{C}$  is required (Kajiyoshi et al. "Preparation of tetragonal barium titanate thin film on titanium metal substrate by hydrothermal method", *J. Am. Ceram. Soc.*, 74, pp 369-74, 1991).

The cubic-phase and lower density of hydrothermally produced barium titanate can be explained because of the existence of lattice defects such as hydroxyls ( $\text{OH}^-$ ) in small particles. According to Vivekanandan et al., the  $\text{OH}^-$  included in barium titanate particles prepared hydrothermally contribute to the stabilization of a cubic phase at room temperature which is normally observed only at temperatures above the ferroelectric Curie point at  $130^\circ\text{C}$  (Vivekanandan et al, Hydrothermal preparation of  $\text{Ba}(\text{Ti,Zr})\text{O}_3$  powders, *Mater. Res. Bull.*, 22, pp 99-108, 1986). In a publication by Hennings et al., the authors indicate that tetragonal phase barium titanate is obtained upon heat-treatment above  $600^\circ\text{C}$  (Hennings et al., Characterization of hydrothermal barium titanate, *J. Eur. Ceram. Soc.*, 9, pp 41-46, 1992).

In a later publication, Hennings et al. (Defect chemistry and microstructure of hydrothermal barium titanate, *J. Am. Ceram. Soc.* 84(1), pp 179-182, 2001) show that

for commercial (Cabot Performance Materials, BT4) and self-prepared cubic-phase barium titanate of initial density  $5.65 \pm 0.05 \text{ g.cm}^{-3}$ , all  $\text{OH}^-$  are removed upon heat treating above  $600^\circ\text{C}$  as showed by thermogravimetric analysis, but only as of  $800^\circ\text{C}$  the density starts increasing reaching  $5.97 \text{ g.cm}^{-3}$  at  $1000^\circ\text{C}$ . Also in the temperature  
5 range  $20\text{-}400^\circ\text{C}$ , there is a large discrepancy between the theoretical density calculated from the lattice parameters ( $5.93 \text{ g.cm}^{-3}$ ) and the pycnometric density ( $5.65 \text{ g.cm}^{-3}$ ). The model involving defects due to the presence of  $\text{OH}^-$  in the structure up to  $600^\circ\text{C}$  however explains the pycnometric density. The lower density of hydrothermal barium titanate is also explained in terms of intragranular porosity  
10 which can only disappear with a heat-treatment above  $800^\circ\text{C}$ .

However heat-treating barium titanate powders to obtain tetragonal phase ( $\sim 600^\circ\text{C}$ ) and a higher density ( $\sim 1000^\circ\text{C}$ ) leads to the common drawbacks of larger particle size, wide particle size distribution and agglomeration.

To be noted that recent work reported that hydrothermal tetragonal barium titanate  
15 can be prepared at a temperature as low as  $240^\circ\text{C}$  over several days of synthesis however obtaining submicrometer-sized, even millimeter-sized particles (in Xu et al. Tetragonal Nanocrystalline Barium Titanate Powder: Preparation, Characterization and Dielectric Properties, J. Am. Ceram. Soc., 86(1) pp 203-205, 2003). As reported by Asiaie et al. (in Characterization of submicron particles of tetragonal  $\text{BaTiO}_3$ , Chem, Mater., 8 pp 226-234, 1996) tetragonal barium titanate  
20 was obtained after 14 days of hydrothermal treatment at  $240^\circ\text{C}$ . As a consequence of the long ageing time, particle growth to  $0.5 \mu\text{m}$ . was observed. Xu et al. (idem, 2003) have reported the preparation of nanometer sized tetragonal barium titanate using the hydrothermal process for 12 hours in a 30 ml vessel; however the TEM  
25 picture shows a wide size distribution. No information about the density of those tetragonal barium titanate powders is reported (above cited Xu et al., Asiaie et al.).

EP 1,205,440 describes the preparation of spherical tetragonal barium titanate particles which have an average primary particle diameter (measured by image analysis from electron micrographs) from  $0.05$  to  $0.5 \mu\text{m}$  and a fairly even particle  
30 size distribution by a process which involves adding an aqueous barium salt solution to a titanium hydroxide solution in the presence of a carboxylic acid to produce starting barium titanate particles by ageing usually at  $70^\circ\text{C}$  for 2 hours. When the ageing time is less than 0.5 hour, a sufficient effect may not be obtained. These particles are then hydrothermally treated to obtain spherical cubic-phase barium  
35 titanate particles which are calcined at temperatures from  $500$  to  $1,200^\circ\text{C}$  to transform the spherical cubic-phase barium titanate particles into spherical tetragonal barium titanate particles.

US 5,445,806 describes the preparation of barium titanate obtained by a wet reaction (including coprecipitation, hydrolysis, hydrothermal or thermal reaction)

from an aqueous solution, followed by calcining at temperature of at least 1000°C for several hours to produce tetragonal barium titanate powder of average primary particle size (measured by image analysis) below 0.3µm, particularly 0.05 to 0.25µm and with uneven shape (see its Fig. 2).

- 5 The Paper "Aqueous Synthesis of Submicron and Nanometre BaTiO<sub>3</sub> Particles in a Segmented Flow Tubular Reactor (SFTR)" Buscaglia et al, CIMTEC 2003 – 10<sup>th</sup> International Ceramics Congress – Part A, describes the production of high quality barium titanate powder using the SFTR technology. The resulting powders show improved quality in comparison to batch synthesis. Fine powders with tailored size  
10 (about 30 nm) could be obtained, that show a certain degree of agglomeration. It has been experimentally verified that these particles are of cubic structure and have high sodium content.

- US 4,764,493 describes the preparation of barium titanate from a suspension or a solution at a temperature below the boiling point of the solvent for at least 48 hours  
15 to obtain barium titanate having a density of at least 5.7 g.cm<sup>-3</sup>, a mean primary particle size of about 0.09-0.2µm with a geometric standard deviation of 30% and a substantially spherical shape.

- US 6,205,015 describes the preparation of dielectric ceramic components and the method for producing such devices using a tetragonal barium titanate powder with  
20 the following characteristics: the amount of hydroxyl groups in the crystal lattice is 0.1 wt. % to about 1 wt. %, the maximum particle size is about 0.5µm or less and the average particle size is about 0.1-0.3µm, individual particles of the barium titanate powder comprise a low-crystallinity portion and a high-crystallinity portion, and the diameter of the low-crystallinity portion is less than about 0.65 times the  
25 particle size of the powder. The related US patent 6,303,529 describes the preparation of dielectric ceramic components using a tetragonal barium titanate powder with the following characteristics: the amount of hydroxyl groups in the crystal lattice is 2.0 wt. % or less the maximum particle size is about 0.3µm or less and the average particle size is about 0.05-0.15µm, individual particles of the  
30 barium titanate powder comprise a low-crystallinity portion and a high-crystallinity portion, and the diameter of the low-crystallinity portion is about 0.5 times or more the particle size of the powder. Such barium titanates can be obtained by thermally treating barium titanate powder which is produced through a wet synthesis method such as a hydrothermal synthesis method, a hydrolysis method, or a sol-gel method.  
35 There may also be employed a solid phase synthesis method in which a carbonate, an oxide, etc of elements constituting the barium titanate powder are mixed and thermally treated. As discussed above, these processes all lead to disadvantages.

EP 1,013,608 describes the chemical preparation of a barium titanate powder from an aqueous solution with an average primary particle size (determined by image

analysis) ranging from 0.1 to 1.0 $\mu$ m with a CV value (standard deviation of the particle size/the average particle size of the particle size distribution) of 21 to 40% (the corresponding CV value of the inventive examples reported below is from 11 to 15%). No indication of the particle shape is provided.

- 5 US 6,485,701, EP 1,264,809 and EP 1,130,004 describe the manufacturing method for an oxide having a perovskite structure by a solid state reaction of a metal oxide powder and a metal carbonate powder, which can be used to produce tetragonal barium titanate with low OH content, with average primary particle diameter (determined by image analysis) in the range 0.03 to 0.2 $\mu$ m. The solid state reaction  
10 requires temperatures above 900°C. This leads to the formation of strong agglomerates necessitating milling which leads to a wide particle size distribution.

US 2003/0044347 and EP 1,415,955 describe the preparation of tetragonal barium titanate and barium titanate based compounds by aqueous reaction up to boiling point followed by crystallization of the BaTiO<sub>3</sub> at 600°C or above for several hours.

- 15 The product's BET specific surface area is between 0.1 m<sup>2</sup>/g and 9.7 m<sup>2</sup>/g corresponding to sizes from about 10 to about 0.1 $\mu$ m.

- US 4,863,883 describes substantially spherical powders of doped barium titanate compositions prepared by hydrothermal synthesis from titanium dioxide and hydroxides at temperatures up to 225°C. Those powders have the following  
20 characteristics: their median primary particle size as determined by image analysis in the range of 0.05 to 0.4 $\mu$ m, the primary particle size distribution, as determined by image analysis, has a quartile ratio (QR=d<sub>75</sub>/d<sub>25</sub>) less than or equal to 1.5, the median primary particle size, as determined by image analysis and by sedimentation, agree within a factor of two, the particle size distribution, as  
25 determined by sedimentation, has a quartile ratio (QR=d<sub>75</sub>/d<sub>25</sub>) less than or equal to 2.0, and the particle size, as determined by image analysis and by surface area, agree within a factor of two. No indication about the density and the crystal structure of the compounds is provided.

- US 4,829,033 describes substantially spherical powders of barium titanate based  
30 dielectric compositions prepared by heating suspensions containing the hydrous tetravalent oxides with selected divalent oxides or hydroxides to which a barium hydroxide solution is added before heating to a higher temperature. Those powders have the following characteristics: their median primary particle size as determined by image analysis in the range of 0.05 to 0.4 $\mu$ m, the primary particle size  
35 distribution, as determined by image analysis, has a quartile ratio (QR=d<sub>75</sub>/d<sub>25</sub>) less than or equal to 1.5, the median primary particle size, as determined by image analysis and by sedimentation, agree within a factor of two, the particle size distribution, as determined by sedimentation, has a quartile ratio (QR=d<sub>75</sub>/d<sub>25</sub>) less

than or equal to 2.0. No indication about the density and the crystal structure of the compounds is provided.

### Summary of the Invention

In a main aspect, the invention provides a method of producing a powder of barium titanate, including solid solutions and doped compounds thereof, which powder is composed of non agglomerated ultrafine particles of tetragonal structure, of high density corresponding to at least 90% of the intrinsic density of a large crystal of the corresponding compound, whose individual particles have an isotropic shape of given aspect ratio. All particles are to be smaller than 1 micron and have a particle size distribution measured by sedimentation characterized by a span  $(d_{v90}-d_{v10})/d_{v50}$  less than 1, where  $d_{v90}$  refers to a value such that 90% of the powder volume is made of smaller sizes,  $d_{v10}$  refers to a value such that 10% of the powder volume is made of smaller sizes, and  $d_{v50}$  refers to the volume median diameter such that 50% of the powder volume is made of smaller sizes, and wherein the ratio of Ba:Ti including their substituents and dopants is very close to ideal stoichiometry.

The method according to the invention for producing powder to this specification comprises two main steps:

(a) a first stage reaction between reactants composed of barium and titanium and optional selected substituents and dopants by preparing at room temperature a liquid-containing reaction mixture and subjecting the reaction mixture to reaction temperatures in the range 80°C, preferably 85°C, to boiling temperature, at a pressure of 1 bar, for a reaction period up to 20 minutes to produce a powder of barium titanate composed of non-agglomerated ultrafine particles having a shape of the above-mentioned given aspect ratio, and cubic structure, of low density corresponding at most to 85% or 90% of said intrinsic density, all particles being smaller than 1 micron and having a particle size distribution span  $(d_{v90}-d_{v10})/d_{v50}$  less than 1, and wherein the ratio of Ba:Ti including their substituents and dopants is very close to ideal stoichiometry; followed by

(b) subjecting the powder of cubic structure produced in step (a) to a second stage solvothermal post treatment at a temperature less than 400°C to convert the particles of low density to ultrafine particles of increased density corresponding to at least 90% of said intrinsic density and convert the isotropic, cubic-structure particles to tetragonal barium titanate particles, while maintaining the size of all particles below 1 micron, the particle size distribution span  $(d_{v90}-d_{v10})/d_{v50}$  below 1, and the ratio of Ba:Ti including their substituents and dopants very close to ideal stoichiometry.

The present invention thus provides a two-step process to prepare high-density tetragonal barium titanate composed of ultrafine particles in the median size range of for example 0.010 to 0.5 microns with a given aspect ratio (isotropic in shape),

usually multi-faceted and preferably predominantly of cube-like shape, narrow particle size distributions, lack of agglomerates, a Ba/Ti ratio, including their substituents and dopants, very close to the ideal stoichiometry (e.g. Ba/Ti ratio very close to 1) and high purity. This process involves the use of low cost raw materials and does not require either subsequent milling or high temperature thermal treatment, though subsequent thermal treatment is possible.

The span of particle size distribution is normally maintained at the same value or can be decreased from step (a) to step (b), with a shift of the particle size distribution ( $d_{v10}$ ,  $d_{v50}$ ,  $d_{v90}$ ) towards higher sizes after the solvothermal post-treatment of step (b). Thus, the particles may grow during the solvothermal post-treatment step (b), and may alter shape, however maintaining essentially the same aspect ratio (ratio of their key dimensions). In particular, the growth of isotropic particles may produce a change of shape from a generally spherical shape to a non-spherical faceted shape, for instance a cube-like shape, however conserving isotropy. The solvothermal post treatment step also produces a change of the crystal structure of the barium titanate particles from cubic to tetragonal crystal structure.

The cubic-structure particles produced in step (a) usually have a generally spherical shape, whereas the tetragonal particles produced in step (b) usually have an isotropic and faceted individual particle shape, in particular a cube-like shape. Such faceted individual particle shapes are isotropic but non-spherical and can be cube-like (six faces) or multifaceted like a hexagonal dipyramid (24 faces), but usually with no more than 30 or 48 faces.

Step (b) of the method can be followed by a heat treatment during which part or all of the particles change shape.

The tetragonal powder produced in step (b) usually has more than 50% and preferably more than 80 or 90% of faceted particles in particular of cube-like shape, however some or all of the particles may lose the faceted or cube-like shape when the particles undergo a subsequent heat treatment. The powder product, produced in step (b), or after a subsequent heat treatment, can for example have 5-95% of individual particles with an isotropic and faceted individual particle shape, in particular a cube-like shape. After various post heat treatments, the amount of the individual particles that retain an isotropic faceted shape in particular a cube-like shape can be 5-80%, 5-50% or 5-20%, for example.

This process can be applied in particular to the preparation of high density barium titanate or solid solutions of barium titanate having the general structure  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$ , where A represents one or more divalent metals such as calcium, lead, strontium, magnesium and zinc and B represents one or more tetravalent metals such as tin, zirconium and hafnium, as well as to high density doped barium titanate and

high density doped solid solutions of barium titanate or other oxides based on barium and titanium having the general structure  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$ .

- Firstly a low temperature (temperature usually in the range 80 or 85°C up to boiling temperature at a pressure of 1 bar) continuous process, which may preferably
- 5 implement Segmented Flow Tubular Reactor technology according to WO98/02237, is used to prepare ultrafine (e.g. spherical shaped particles of median particle size 0.005 – 0.5 microns) particles of low-density cubic-phase barium titanate exhibiting the above-described final characteristics except high density, the final tetragonal structure and the final usually faceted shape such as cube-like shape. In a Segmented
- 10 Flow Tubular Reactor (SFTR) reaction volumes, conveniently of up to 100ml, usually less than 20ml, are separated by a segmenting fluid that is not miscible with the reaction mixture, in a tube typically having a diameter of about 2-10 mm. A reaction time of few minutes already is long enough to produce these low density barium titanate powders. The reaction time is the time necessary to form the barium titanate.
- 15 In a tubular reactor such as an SFTR it can be convenient to maintain the reacted mixture at or about the reaction temperature for the time necessary to exit the reactor, usually a few minutes.

- Alternatively, step (a) is performed using a continuous tubular reactor, a discontinuous fed-batch reactor of preferably up to 1 liter reaction volume, or using a
- 20 continuous mixed suspension-mixed product removal (MSMPR) reactor of preferably up to 1 litre residence volume, taking care to rapidly heat the reaction mixture from room temperature to the reaction temperature, usually in less than 3 minutes and if necessary with stirring. Typically, a volume of up to 500 ml, in particular up to 200 ml, of the reaction mixture of step (a) is heated to the reaction temperature in a
- 25 time of less than 1 minute.

- Secondly, a high temperature (temperature usually above 100°C, and preferably in the range 200-350°C) solvothermal post-treatment is applied to the e.g. spherical-shaped cubic-phase barium titanate particles produced in the first step to transform them into high-density particles exhibiting all the above-described final characteristics. For
- 30 barium titanate, the transformation from the cubic-phase low density to high density and tetragonal is achieved in only 3 or 6 hours, depending on the other chemical and physical parameters, likewise for most barium titanate based solid solutions and doped compounds thereof, where however the formation of faceted shapes other than cube-like is possible.
- 35 Solvothermal treatment relates to the treatment of a solution, suspension, dispersion, emulsion or gel using a solvent at a pressure and temperature above the normal pressure and boiling point of said solvent. If water is the solvent used then such a treatment is called a hydrothermal treatment and is carried out at a pressure above 1 bar and temperature above 100 °C.

This two-step process allows for the production of commercial quantities of high quality barium titanate particles including solid solutions and doped compounds thereof for specific applications, e.g. adapted for small size multilayer ceramic capacitors and embedded capacitances requiring the smallest barium titanate particles.

- 5 The invention also relates to the barium titanate powder, including solid solutions and doped compounds thereof, produced by the inventive method, in particular where the resulting particles have the defined characteristics and where at least 50%, preferably at least 80 or 90% of the individual particles have an isotropic faceted shape, in particular a cube-like shape.
- 10 Specifically, the invention includes a powder of barium titanate  $\text{BaTiO}_3$  including solid solutions and doped compounds thereof, that is obtainable by the inventive method, which powder is composed of ultrafine particles, wherein no agglomerates are bigger than  $1\text{ }\mu\text{m}$ , preferably no bigger than  $0.8\text{ }\mu\text{m}$ , and having the following characteristics:
  - 15 a) a primary particle size distribution determined from image analysis with a median number diameter  $d_{n50}$  comprised in the range  $0.005$  to  $0.250\text{ }\mu\text{m}$ , a maximum primary particle size of  $0.5\text{ }\mu\text{m}$ , a number size distribution span  $(d_{n90}-d_{n10})/d_{n50}$  of  $0.8$  and below, and a geometric standard deviation  $\sigma_g (=d_{n50}/d_{n84})$  from  $0.75$  to  $<1$ ,
  - 20 b) a particle size distribution measured by sedimentation with a median volume diameter  $d_{v50}$  comprised in the range  $0.01$  to  $0.5\text{ }\mu\text{m}$ , without any hard agglomerate bigger than  $0.8\text{ }\mu\text{m}$ , a volume size distribution span  $(d_{v90}-d_{v10})/d_{v50}$  less than  $0.95$  and a geometric standard deviation  $\sigma_g (=d_{v50}/d_{v84})$  from  $0.70$  to  $<1$ ,
  - c) a factor of agglomeration ( $F_{AG} = d_{v50}/d_{BET}$ ) smaller than  $2.1$ , preferably smaller than  $1.6$ ,
  - 25 d) a ratio between the sedimentation  $d_{v50}$  and the image analysis  $d_{v50}$  in the range  $1.0$  to  $1.5$ ,
  - e) a tetragonal structure,
  - f) a density of at least  $90\%$  of the intrinsic density of a large crystal of the corresponding compound,
  - 30 g) a hydroxyl content less than  $1\text{ wt}\%$ , and
  - h) a ratio of Ba:Ti including their substituents and dopants equal to  $1.00$  plus or minus  $1$  atomic percent.

Further features and embodiments of the invention are set out in the claims.

#### Brief Description of the Drawings

- 35 Fig.1 is an X-ray diffraction pattern of cubic barium titanate obtained Example 1.

Fig. 2 is the volumetric particle size distribution of the barium titanate obtained in Example 1.

Fig. 3 is an SEM micrograph of the spherical-shaped barium titanate obtained in Example 1.

5 Fig. 4 is the particle size distribution of the barium titanate obtained in Example 4.

Fig. 5 is an SEM micrograph of the barium titanate obtained in Example 4.

Fig. 6 is an X-ray diffraction pattern in the region 2-theta 52 to 54 degrees of tetragonal barium titanate obtained in Example 4.

10 Fig. 7 is an X-ray diffraction pattern in the region 2-theta 52 to 54 degrees of tetragonal barium titanate obtained in Example 8.

Fig. 8 is the particle size distribution of the barium titanate obtained in Example 8.

Fig. 9 is an SEM micrograph of the barium titanate obtained in Example 8.

Fig. 10 is the particle size distribution of the barium titanate obtained in comparative Example 9.

15 Fig. 11 is an SEM micrograph of the barium titanate obtained in comparative Example 9.

Fig. 12 is the particle size distribution of the barium titanate obtained by a comparative post heat treatment in Example 10.

#### Detailed Description of the Invention

##### 20 **Production of Basic Compounds**

The present invention provides a two-step process to prepare high-density perovskite tetragonal barium titanate powders composed of ultrafine particles in the median size range 0.01 to 0.5 microns with isotropic shape, narrow particle size distributions, lack of agglomerates, a Ba/Ti ratio very close to 1 and high purity. This process involves  
25 the use of low cost raw materials and does not require either subsequent milling or high temperature thermal treatment, though subsequent thermal treatments are possible.

The first step is a low temperature continuous process where either a discontinuous one liter fed-batch reactor; or a continuous mixed suspension-mixed product removal  
30 (MSMPR) reactor of one liter residence volume; or any other process generating continuously small reaction volumes, or a continuous tubular reactor or preferably a Segmented Flow Tubular Reactor (SFTR) as described in WO98/02237, is used to prepare ultrafine low-density cubic barium titanate particles.

35 A first aqueous or non-aqueous solution is prepared using soluble organic or inorganic barium and titanium reactants (e.g. BaCl<sub>2</sub> and TiCl<sub>4</sub>) with Ba/Ti ratio larger

than 1 and titanium concentration in the range about 2.0-0.01 mol/Kg. The Ba/Ti ratio changes as a function of titanium concentration, in the range 1.01-1.2 to lead to a cubic barium titanate product of ideal stoichiometry. A second aqueous or non-aqueous solution is prepared from a strong base compound (e.g. NaOH) or another  
5 base compound such as ammonia, or tetramethylammonium hydroxide (TMAH). If the barium source is barium hydroxide, the use of a mineralizer may not be strictly required to ensure barium titanate crystallization. The two solutions are mixed, optionally using micromixers, and mechanically homogenized in order to produce a reaction mixture in the form of a liquid solution, suspension, dispersion, emulsion or  
10 a gel with a pH which is high enough, e.g. 13, preferably in the range of about 13.5-14 after complete transformation into barium titanate. Such as-prepared gel is an amorphous titanium hydroxide network with adsorbed barium ions. The reaction mixture thus prepared in the form of a liquid solution, dispersion, emulsion, suspension, or gel at room temperature is progressively fed to a SFTR and segmented  
15 with an immiscible fluid. The SFTR is maintained at the appropriate reaction temperature (in the range from 80 or 85°C up to boiling temperature at a pressure of 1 bar), the reaction mixture being brought rapidly (in less than 30 seconds) from room temperature to the reaction temperature, for successive volumes of the reaction mixture that are less than 200 ml (the individual segmented reaction volumes can be  
20 about 20-100 ml). The reaction time is typically in the range 2 to 20 minutes. During the reaction time ("rapid ageing"), the gel or reaction mixture is transformed into single-phase low-density suspension of generally spherical-shaped barium titanate particles. The suspension is recovered and the powder washed in water until pH is above 11 and then in an NH<sub>4</sub>OH solution until chlorine concentration in the  
25 supernatant is below 10ppm.

X-Ray Diffraction analysis detects a typical cubic-phase barium titanate pattern without any secondary phases such as barium carbonate, Ba-rich or Ti-rich mixed oxides. The Ba/Ti ratio in the powder has been measured by Inductively Coupled Plasma spectrometry and is equal to  $1.00 \pm 0.01$ . The density measured by helium  
30 pycnometry was usually in the range  $5.0\text{-}5.1\text{ g.cm}^{-3}$ , this value is also consistent with a cubic-phase low-density barium titanate. Thermogravimetric analysis carried out from 25 to 500°C detects a weight loss of about 4 % attributed to the removal of hydroxyls present within the barium titanate structure; this is also consistent with a cubic-phase low-density barium titanate.

35 Adjusting the synthesis parameters such as the reactant concentration, Ba/Ti ratio in the first solution, the temperature and the reaction mixture or gel flow rate allows for fine-tuning the median volume diameter ( $d_{v50}$ ) of the barium titanate particles in the range of 0.005 to 0.5 microns. The particle size distribution is measured using a sedimentation technique. The width of the volume particle size distribution is

evaluated by the span, which corresponds to  $(d_{v90}-d_{v10})/d_{v50}$ . This span is below 1.0, typically about 0.7 ( $d_{vX}$  is a particle diameter meaning that X% of the powder volume is made of smaller sizes). At this stage, the particles have a generally spherical shape.

5 The second step is a high temperature solvothermal post-treatment applied to the cubic-phase low-density barium titanate particles produced in the first step to transform them into high-density tetragonal barium titanate. The temperature is above 100°C, preferably in the range 200-350°C.

10 During the second step, the generally spherical low-density barium titanate particles of cubic crystal structure are transformed into high-density, generally cube-shaped, tetragonal barium titanate particles, though other shapes such as non-spherical faceted shapes are possible.

15 Preferably, step (b) is performed by adding the barium titanate powder produced in step (a) to an aqueous solution of at least one metal compound at a pH above 9 or to a non-aqueous solution of at least one metal compound to produce a suspension, and subjecting the suspension to a solvothermal post treatment at a pressure above 1 bar and a temperature between 100°C and 400°C, preferably in the range 200-350°C, for a residence period of 2 to 20 hours.

20 An aqueous or non aqueous barium solution at high pH is prepared e.g. from a barium salt or alkoxide and a strong base reactant (e.g.  $\text{BaCl}_2$  and  $\text{NaOH}$  or  $\text{NH}_4\text{OH}$  or TMAH) or using  $\text{Ba}(\text{OH})_2$ . Alternatively a non-aqueous solvent is used. The solution is mixed with the cubic-phase barium titanate particles prepared in the first step using the SFTR and the resulting suspension is homogenized. The solid/liquid ratio can be as large as 50% weight. The suspension is transferred to an autoclave of any suitable volume and treated at a temperature above 200°C and for times from 2 to 20 hours  
25 (also, for comparative purposes, 1 hour). The second step can also be carried out in a continuous solvothermal apparatus. The suspension is recovered and the powder washed in water until pH is above 11 and then in an  $\text{NH}_4\text{OH}$  solution until chlorine concentration in the supernatant is below 10ppm. After washing, the powder is dried. The pressure applied in the solvothermal is in the range above 1 to 100 bars,  
30 preferably in the range 5 to 75 bars. The heating of the solvothermal mixture can be ensured using various devices, e.g. a furnace, a bath or microwaves. Several temperature steps may be applied within the same solvothermal post-treatment. Also several solvothermal post-treatments may be applied with washing of the barium titanate powder and modification of the chemical composition of the above described  
35 aqueous or non-aqueous solution.

X-Ray Diffraction analysis detects a typical tetragonal barium titanate pattern without any secondary phases such as barium carbonate, Ba-rich or Ti-rich mixed oxides. The Ba/Ti ratio in the powder has been measured by Inductively Coupled Plasma

spectrometry (ICP-AES) and is equal to  $1.00 \pm 0.01$ . The density measured by helium pycnometry is usually in the range  $5.60\text{--}5.75\text{g.cm}^{-3}$  but may be increased up to  $5.9\text{g.cm}^{-3}$ . Thermogravimetric analysis carried out from 25 to  $500^{\circ}\text{C}$  detects a weight loss of only 0.80% maximum attributed to the removal of residual hydroxyls present within the barium titanate structure; this value slightly varies upon the post-treatment conditions.

Adjusting the synthesis parameters such as the reactant concentration, Ba/Ti ratio in the first solution, the temperature and the post-treatment time allows for fine-tuning the volume median diameter ( $d_{v50}$ ) of the barium titanate particles in the range of 0.01 to 0.5 microns. The particle size distribution is measured using a sedimentation technique. The width of the size distribution is evaluated by the span, which corresponds to  $(d_{v90}-d_{v10})/d_{v50}$ . This span is in the range 0.4-0.8, depending on the post-treatment parameters. The  $F_{AG}$ , determined from the volume median diameter ( $d_{v50}$ ) obtained by sedimentation and from the BET diameter  $d_{BET}$  (calculated from the specific surface area and density), is in the range 1.6 to 2.1. Determination of volume particle size distributions has also been carried out from image analysis in order to compare the ratio between the volume sizes  $d_{v10}$ ,  $d_{v50}$ , and  $d_{v90}$  determined from sedimentation and from image analysis. Those ratios are usually in the range 1.0 to 1.6, ensuring that the powders consist of mostly soft agglomerates that can be easily dispersed. Soft agglomerates correspond to several particles held together by weak cohesive forces (van der Waals, hydrophobic or electrostatic forces) which can be broken by milling or sonication.

This solvothermal post-treatment is necessary to transform the high quality, generally spherical-shaped, cubic-phase barium titanate prepared in the first step of the process into high quality, generally cube-like or facet shaped, tetragonal barium titanate while maintaining the other key characteristics obtained through the first step (isotropic, narrow particle size distributions, lack of hard agglomerates and a Ba/Ti ratio very close to 1). Hard agglomerates are cohesive masses consisting of particulate subunits which cannot be broken by sonication or soft milling. The duration of the solvothermal post-treatment also allows for increasing the volume median particle size ( $d_{v50}$ ) while ensuring the particle size distribution is still narrow. In addition, after performing the solvothermal post-treatment it has been shown that the powder produced using sodium containing raw materials contains about 5 to 10 times, in some cases 10 to 20 times less sodium than before, with a sodium concentration in the range 40 to 100ppm, in some cases 10 to 35ppm depending on the post-treatment conditions.

Using the barium titanate powder prepared according to the method to prepare a ceramic body allows for determining dielectric properties of such a ceramic body. Cylindrical rods of 8-9.5 mm in diameter are made by isostatic pressing and sintered

at different temperatures for a fixed time. Discs approximately 1 mm thick were cut from the rods, flat parallel faces were ground with emery and gold electrodes were evaporated onto them in order to prepare a capacitor. The capacitance and dielectric losses were measured between 30 and 170°C and were satisfactory for commercial applications

#### Optional Heat Treatments

The high quality cube-like tetragonal barium titanate powder may further be heat treated in a continuous or discontinuous furnace at a temperature typically within the range from 400 to 1200°C for durations usually from a few minutes up to 48 hours where the heating rate, the type of gas or gas mixtures used as well as their partial pressure may be adjusted. This heat treatment may lead to bigger particles or not. Conditions are selected so that high quality tetragonal barium titanate powder with narrow particle size distribution (span < 1.0), without agglomerates above 0.8 micron is obtained. Conditions can be selected so that the  $F_{AG}$  remains in the same range as before the heat treatment (usually below 2.0) as well as the ratios of the volume sizes obtained by sedimentation and image analysis (below 2.0). Median size may increase as a function of temperature and duration of the heat treatment. Thermogravimetric analysis shows a weight loss attributed to the presence of hydroxyls in the barium titanate structure below 0.7% after such heat treatments. During this optional heat treatment the particles maintain isotropy but lose partially the distinctive multifaceted or cube-like form above 900 to 1200°C, also depending on the other parameters (e.g. time, heating rate, atmosphere). Subsequent milling treatment is possible.

#### Substitution and doping

The electrical properties of barium titanate can be substantially modified by the incorporation of barium and titanium substituents to form solid solutions and/or as dopants.

Solid solutions of barium titanate have the general structure  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$ , where A represents one or more divalent metals such as calcium, cadmium, europium, magnesium, lead, radium, strontium and zinc and B represents one or more tetravalent metals such as cerium, cobalt, iron, hafnium, molybdenum, lead, praseodymium, plutonium, ruthenium, tin, thorium, titanium, uranium, vanadium and zirconium.

One type of barium titanate-based composition has the structure  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$ , where the mole fractions x and y can be in the range of 0 to 1, preferably larger than 0.03, where A represents one or more divalent metal other than barium such as calcium, cadmium, europium, magnesium, lead, radium, strontium and zinc and B represents one or more tetravalent metals other than titanium such as cerium, cobalt,

iron, hafnium, molybdenum, lead, praseodymium, plutonium, ruthenium, tin, thorium, uranium, vanadium and zirconium.

Where the divalent or tetravalent metals are present as dopants, the value of the mole fraction  $x$  and  $y$  may be small, for example less than 0.1.

- 5 In other cases, the divalent or tetravalent metals may be introduced at higher levels to provide a significantly identifiable compound such as barium-calcium titanate, barium-strontium titanate, barium titanate-zirconate and the like.

The powder composition also shows a mole ratio of  $(Ba+A)/(Ti+B)$  close to 1.00.

- 10 Different cations (e.g. calcium, cadmium, europium, magnesium, lead, radium, zinc, cerium, iron, hafnium, molybdenum, praseodymium, plutonium, ruthenium, tin, thorium, uranium and zirconium e.g. in the form of chlorides but not limited to chlorides), one or more as desired, can substitute the barium (or A) and/or titanium (or B) reactants of the first solution used to prepare the reaction mixture in the form of a liquid, suspension, emulsion or gel before completing the first step of the process.
- 15 The low density barium titanate-based particles prepared though the first step of the process are then transformed into high density barium titanate-based particles  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  through the above-described solvothermal post-treatment.

Dopants cover a wide range of metal oxides. These, in general but not limited, represent less than 0.05 mole fraction of the total  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  formulation.

- 20 The dopant or dopants employed may be completely or partially miscible in the perovskite lattice or may be immiscible in the lattice. The dopant is selected from the group consisting of elements such as lanthanum, rare earth elements, lithium, magnesium, yttrium, niobium, tantalum, gallium, molybdenum, tungsten, manganese, copper, iron, cobalt, nickel, chromium, zinc, aluminium, silicon, antimony, lead,
- 25 bismuth, boron and mixtures thereof and is different to A and B.

- It is possible to obtain homogeneously doped low density barium titanate or doped low density  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  compound using a modified procedure of the first step of the process. The desired cations are added in the form of organic or inorganic compounds (e.g. chlorides) to the first aqueous or non-aqueous solution containing soluble barium and titanium reactants either before or after the reaction mixture in the form of liquid, suspension, emulsion or gel is prepared, and then reacted in the SFTR or another type of reactor as described for the manufacture of pure barium titanate. The doped barium titanate or low density doped  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  compound prepared though the first step of the process is then transformed into high density doped barium titanate or high density doped  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  compound through the above-described solvothermal post-treatment.
- 30
- 35

Another route to prepare heterogeneously doped barium titanate with a concentration gradient decreasing from the particles surface to their core is to use pure low density

barium titanate prepared through the first step of the process and modify the second step of the process by replacing (or adding to) the barium compound in the barium solution the desired metallic cation(s) to obtain doped high density barium titanate. Adjusting the dopant concentration in the slightly acidic or alkaline solution (depending on the nature of the dopant element) allows for either getting a gradient or a homogenous distribution of dopant within the particles.

This last route can also be used to prepare doped  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  compound with a concentration gradient decreasing from the particles surface to their core. This route uses low density  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$  compound prepared through the first step of the process and the second step of the process is modified by replacing the barium and/or the A compound in the slightly acidic or alkaline solution by the desired metallic cation(s) in order to produce high density doped  $Ba_{(1-x)}A_xTi_{(1-y)}B_yO_3$ . Adjusting the dopant concentration in the slightly acidic or alkaline solution allows for either producing a gradient or a homogenous distribution of dopant within the particles.

Several solvothermal post-treatments may be carried out one after the other, each one allowing doping with a single or several dopants. Also, a solvothermal post-treatment may be carried out without the presence of a dopant but only with appropriate solvent or solvent mixture. For substituted and doped barium titanate, the particles may adopt other faceted shapes than cube-like after the solvothermal treatment, the faceted shape depending on the substituents/dopants and the process conditions.

#### EXAMPLES

The present invention will be described in more detail hereinunder by Examples, but the Examples are only illustrative and, therefore, not intended to limit the scope of the present invention thereto.

Various properties were measured by the following methods.

(a) The shape of the barium titanate particles was determined from micrographs made during scanning electron microscope (SEM, PHILIPS model XL-30 FEG or 515) observations. The aspect ratio of a particle can also be determined from micrographs. The aspect ratio corresponds to the ratio between the longest dimension and the shortest dimension of the particle, i.e. the aspect ratio of a rod is bigger than 1, the aspect ratio of a disc or platelet is smaller than 1 while it equals 1 for a sphere. More generally, the aspect ratio is defined as the ratio between the particle axis  $h$  and the axis  $b$ . In a Cartesian coordinate system,  $h$  corresponds to the  $z$  direction and  $b$  to the  $x$ - $y$  direction for symmetric bodies. Oblate spheroids will then have aspect ratios  $h/b < 1$  and prolate spheroids will have axial ratios  $h/b > 1$ .

(b) The number primary particle size distributions of the barium titanate particles were measured by image analysis from scanning electron micrograph in order to determine the edges of the cube-shaped primary particles. Two or more touching particles were visually disaggregated and the sizes of the individual primary particles were measured. The edges were used to compute the cumulative number percent distribution as a function of primary particle size. The number particle size distribution is characterized by  $d_{n10}$ ,  $d_{n25}$ ,  $d_{n50}$ ,  $d_{n75}$ ,  $d_{n84}$ ,  $d_{n90}$ :  $d_{nX}$  is a particle diameter meaning that X% of the number of particles is made of smaller sizes e.g. 10% of the particles is composed of smaller particle sizes for  $d_{n10}$ . Also  $d_{n100}$  corresponds to the biggest particle edge of the number particle size distribution. The width of the number distribution is defined by the span  $[(d_{n90}-d_{n10})/d_{n50}]$ . The smaller the span, the narrower is the particle size distribution. The geometric standard deviation  $\sigma_g$  is calculated from the following ratio  $d_{n50}/d_{n84}$ , it provides another indication of the width of the number size distribution. The closer to 1, the narrower the size distribution is. The quartile ratio, QR, defined as the upper quartile diameter ( $d_{n75}$ ) divided by the lower quartile diameter ( $d_{n25}$ ), is another measure of the width of the distribution: the size distribution is extremely narrow when QR is close to 1, the bigger QR is, the wider is the size distribution.

The equivalent spherical diameters of the cube-shaped particles were used to compute the cumulative volume percent distribution as a function of primary particle size. The volume particle size distribution is characterized by  $d_{v10}$ ,  $d_{v25}$ ,  $d_{v50}$ ,  $d_{v75}$ ,  $d_{v84}$ ,  $d_{v90}$  ( $d_{vX}$  is a particle diameter meaning that X% of the powder volume is made of smaller sizes). The width of the volume distribution is defined by the span  $[(d_{v90}-d_{v10})/d_{v50}]$ . The smaller the span, the narrower is the particle size distribution. The geometric standard deviation  $\sigma_g$  is calculated from the following ratio  $d_{v50}/d_{v84}$ , it provides another indication of the width of the volume size distribution; the closer to 1, the narrower the size distribution is. The quartile ratio, QR, defined as the upper quartile diameter ( $d_{v75}$ ) divided by the lower quartile diameter ( $d_{v25}$ ), is another measure of the width of the distribution: the size distribution is extremely narrow when QR is close to 1, the bigger QR is, the wider is the size distribution.

(c) The volumetric particle size distributions of the barium titanate particles were measured by sedimentation (XDC, BI-XDCP Particle Sizer, Brookhaven Instrument Corporation). The sedimentation process reports particle size distribution of the powder in terms of the equivalent spherical diameter. This measure represents the diameter of a sphere that would settle at the same

- terminal velocity as the particle in question. This method is very representative of the system analyzed if the particles are approximately equiaxed. Also, the particles measured using sedimentation correspond to secondary particles made of several primary particles bound together but acting as single entities. To prepare a suspension for particle size distribution measurement, 0.5g of dry powder was dispersed in 40 ml of a 0.1wt% polyacrylic acid solution. The suspension was sonicated with an ultrasound horn for 15 minutes. The measurement was carried out at 25°C with a disk speed of 1500 rpm (higher disc speeds are used for smaller particles). The particle size distribution is characterized by  $d_{v10}$ ,  $d_{v50}$ ,  $d_{v84}$ ,  $d_{v90}$  :  $d_{v10}$  means that 10% of the powder volume is composed of smaller particle sizes,  $d_{v50}$  means that 50% of the powder volume is composed of smaller particle sizes,  $d_{v84}$  means that 84% of the powder volume is composed of smaller particle sizes and  $d_{v90}$  means that 90% of the powder volume is made of smaller sizes. Also  $d_{v100}$  corresponds to the biggest equivalent diameter of the volume particle size distribution. The width of the volume distribution obtained by sedimentation is defined by the span ( $[d_{v90}-d_{v10}]/d_{v50}$ ). The smaller the span, the narrower is the particle size distribution. The geometric standard deviation  $\sigma_g$  is calculated from the ratio  $d_{v50}/d_{v84}$ , it provides another indication of the width of the volume size distribution. The closer to 1, the narrower the size distribution is. The percentage of agglomerates bigger than 1 micron was also calculated from the particle size distribution data.
- (d) Since particle sizes determined by image analysis and by sedimentation depend on different principles, an exact correspondence in size by these two methods is rarely obtained. Moreover in image analysis touching primary particles are visually disaggregated. In the sedimentation process bound or flocculated particles act as single entities. These entities usually named secondary particles arise because of the existence of some bonding (e.g., necking) between the primary particles giving hard agglomerates which cannot be readily broken down during the sonication process. The ratios between the  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  values determined by sedimentation and those respectively found from image analysis were calculated: the smaller they are, the less primary particles compose a secondary particle determined by sedimentation. As expected, they were usually larger than 1 but smaller than 2.
- (e) The specific surface area value was measured by the BET method using a Gemini Micromeritics apparatus after the powder had been heated for 1 hour at 220°C.

- (f) The densities were measured using helium pycnometry (Micromeritics AccuPyc1330) pycnometry after the powder was heated for minimum 24h at 110°C to make sure all adsorbed water had been eliminated.
- (g) The BET diameter  $d_{\text{BET}}$  is calculated from the BET specific surface area and the density of the powder measured by pycnometry using the following equation  $d_{\text{BET}} = 6 / (\text{density} \times \text{BET specific surface area})$ , assuming particles are spherical or cube-shaped and the primary particle size distribution is narrow. Dispersibility of the powders was assessed by comparing the  $d_{\text{BET}}$  and median size  $d_{v50}$  determined by sedimentation procedure. The agglomeration factor ( $F_{\text{AG}} = d_{v50} / d_{\text{BET}}$ ) indicates the state of agglomeration of the powder under investigation.
- (h) The Ba/Ti ratio was measured using inductively coupled plasma spectrometry (ICP-AES, Perkin Elmer, Plasma 2000 software) as well as the sodium concentration in the powder.
- (i) Phase purity was investigated by X-ray diffraction (XRD, Co  $K_{\alpha}$  radiation, model PW1710, Philips, Eindhoven, The Netherlands) on as-prepared powders and after calcination for 6 h at 950°C and 1350°C. Phase composition after annealing is very sensitive to the overall Ba/Ti molar ratio of the solid. Formation of secondary phases is already observed by XRD for deviations of  $\pm 1$  at.% from the ideal stoichiometry if the diffraction pattern is carefully collected. As to the tetragonality of the barium titanate particles of the present invention, the value of  $c/a$  has been calculated, wherein 'a' and 'c' are lattice constants for the respective 'a' and 'c' crystal axes.
- (j) Thermogravimetric analysis (Mettler TG50 thermobalance) measured in the range 200 to 500°C allowed for determining the amount of hydroxyls contained in the barium titanate structure before and after the solvothermal post-treatment. The measurements were carried out after the powder was heated for minimum 24h at 110°C to make sure all adsorbed water had been eliminated.

### 30 Production of Powder from First Step of the Process

#### Example 1a

To prepare 1 kg of cubic-phase barium titanate with a volume median size ( $d_{v50}$ ) of 75nm determined by sedimentation, twelve litres of gel were prepared at ambient temperature by mixing 6 litres of a NaOH solution at a concentration of 4.42 mol/kg and 6 litres of BaCl<sub>2</sub> and TiCl<sub>4</sub> solution at a concentration of 0.561 mol/kg and 0.550 mol/kg respectively. The gel was fed to a Segmented Flow Tubular Reactor comprising a tube immersed in a water bath maintained at 98°C. A heating time from room temperature to 98°C within less than 20 seconds was ensured. The residence

time of the gel in the tube was between 3 to 5 minutes, over the 2 minutes reaction time required to transform the gel into a suspension of barium titanate nanoparticles at this temperature. The suspension was collected at the outlet of the tube in a vessel maintained at room temperature. The suspension was then allowed to decant for 2 hours and about 10 litres of supernatant were removed from the storage vessel. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1 and illustrated by Figure 2, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 75nm with a  $d_{v10}$  and  $d_{v90}$  of 48nm and 116nm respectively. Figure 1 shows the X-ray diffraction pattern of the resulting cubic-phase barium titanate, also since no diffraction peaks attributed to substances other than  $BaTiO_3$  were recognized, it was confirmed that the obtained particles consisted of  $BaTiO_3$ . The powder is considered as cubic-phase as no peak splitting has been identified and  $c/a = 1.000$  (see Table 3). Also Figure 3 shows a SEM micrograph of the spherical-shaped barium titanate particles prepared. The density of this powder is  $5.00 \text{ g.cm}^{-3}$ . The Ba/Ti equals 1.00. The powder also contains 500ppm of sodium inherent to NaOH used in this synthesis route. Finally, thermogravimetric analysis shows a weight loss of 4.2% in the range 200 to 500°C attributed to the presence of hydroxyls in the barium titanate structure.

#### 20 Example 1b

To prepare 1 kg of cubic-phase barium titanate with a volume median size ( $d_{v50}$ ) of 155nm determined by sedimentation, in a variation of Example 1a, twelve litres of gel were prepared at ambient temperature by mixing 6 litres of a NaOH solution at a concentration of 1.6 mol/kg and 6 litres of  $BaCl_2$  and  $TiCl_4$  solution at a concentration of 0.22 mol/kg and 0.20 mol/kg respectively. The gel was fed to a Segmented Flow Tubular Reactor comprising a tube immersed in a water bath maintained at 87°C. A heating time from room temperature to 87°C within less than 20 seconds was ensured. The residence time of the gel in the tube was between 8 to 10 minutes, over the time (about 3-4 minutes) required to transform the gel into a suspension of barium titanate nanoparticles at this temperature. The suspension was collected at the outlet of the tube in a vessel maintained at room temperature. The suspension was then allowed to decant for 2 hours and about 10 litres of supernatant were removed from the storage vessel. The powder was washed in several steps and then dried. The powder was then characterized.

35 The volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 155nm with a  $d_{v10}$  and  $d_{v90}$  of 101nm and 252nm respectively. X-ray diffraction pattern shows cubic barium titanate and no diffraction peaks attributed to substances other than  $BaTiO_3$  were recognized. The powder is considered as cubic-phase as no peak splitting has been identified and  $c/a = 1.000$  (see Table 3). SEM micrograph

observations showed spherical-shaped barium titanate particles. The density of this powder is  $5.00 \text{ g.cm}^{-3}$ . The Ba/Ti ratio equals 1.00.

The resulting particles can then be subjected to hydrothermal post treatment according to the invention, for instance as described in Examples 2 to 8 for Example 1a's particles.

### Second Step of the Process

#### Example 2 - Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out by mixing 2g of barium titanate produced according to procedure described in Example 1a with 30g of a  $\text{Ba(OH)}_2$  solution. This solution was initially prepared by dissolving 16.7g of  $\text{Ba(OH)}_2$  in 1 litre of water. The suspension was placed in a 40 ml hydrothermal autoclave, itself placed in an oven heated at  $250^\circ\text{C}$ . Reaching a temperature of  $250^\circ\text{C}$  inside the autoclave required 1 hour, the autoclave was then allowed to stay at this temperature for 1 hour and then to cool naturally to room temperature. The suspension was recovered and the supernatant was removed. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 106nm with a  $d_{v10}$  and  $d_{v90}$  of 71nm and 151nm respectively without any agglomerate above 250nm. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is about 1. The particle shape is no longer spherical but predominantly cube-like (85%). The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.002$  (see Table 3). The density of this powder has increased compared to the Example 1a powder: it is  $5.60 (\pm 0.09) \text{ g.cm}^{-3}$ . The powder only contains 34ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.80 % in the range 200 to  $500^\circ\text{C}$  attributed to the presence of residual hydroxyls in the barium titanate structure.

Scaling up of Examples 2 to 8 to a 2-litre autoclave gives the same results.

#### Example 3 - Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out using the same procedure as Example 2, except the hydrothermal autoclave was maintained at  $250^\circ\text{C}$  for 2 hours instead of 1 hour.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 130nm with a  $d_{v10}$  and  $d_{v90}$  of 85nm and 189nm respectively without any agglomerate above 365nm. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and

image analysis is about 1. As shown in Figure 6, the powder is of tetragonal crystal structure as a peak splitting has been identified and  $c/a = 1.006$  (see Table 3). Also according to SEM observations of the barium titanate particles, their shape has now become more cube-like than spherical: more than 85% of the particles have an individual cube-like shape. The density of this powder is  $5.58 (\pm 0.10) \text{ g.cm}^{-3}$ . The powder also only contains 26ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.70% in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

#### Example 4 – Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out using the same procedure as Example 2, except the hydrothermal autoclave was maintained at 250°C for 6 hours instead of 1 hour.

As presented in Table 1 and illustrated by Figure 4, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 187nm with a  $d_{v10}$  and  $d_{v90}$  of 152nm and 229nm respectively without any agglomerate above 265nm. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is about 1.5. Figure 5 shows a SEM micrograph of the particles where 90% of individual particles present a cube-like shape. The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.007$  (see Table 3). The density of this powder has increased compared to the Example 1a powder: it is  $5.73 (\pm 0.06) \text{ g.cm}^{-3}$  and its specific surface area is  $11.6 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti equals 1.00. The powder only contains 31ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.70% in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

#### Example 5 – Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out by mixing 2g of barium titanate produced according to procedure described in Example 1a with 30g of a  $\text{BaCl}_2$  and NaOH solution. This solution was initially prepared by dissolving 13.66g of  $\text{BaCl}_2$  and 5g of NaOH in 1 litre of water. The suspension was placed in a 40ml hydrothermal autoclave, itself placed in an oven heated at 250°C. Reaching a temperature of 250°C inside the autoclave required 1 hour, the autoclave was then allowed to stay at this temperature for 6 hours and then to cool naturally to room temperature. The suspension was recovered and the supernatant was removed. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 154nm with a  $d_{v10}$  and  $d_{v90}$  of 105nm and 217nm

respectively without any agglomerate above 350nm. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is in the range 1.1 to 1.6. According to SEM observations 86% of individual particles present a cube-like shape. The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.006$  (see Table 3). The density of this powder has increased compared to the Example 1a powder: it is  $5.71 (\pm 0.01) \text{ g.cm}^{-3}$  and its specific surface area is  $12.1 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti equals 1.00. The powder only contains 65ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.80 % in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

#### Example 6 – Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out by mixing 2g of barium titanate produced according to procedure described in Example 1a with 30g of a  $\text{BaCl}_2$  and NaOH solution at pH14. This solution was initially prepared by dissolving 13.66g of  $\text{BaCl}_2$  and 40g of NaOH in 1 litre of water. The suspension was placed in a 40ml hydrothermal autoclave, itself placed in an oven heated at 250°C. Reaching a temperature of 250°C inside the autoclave required 1 hour, the autoclave was then allowed to stay at this temperature for 6 hours and then to cool naturally to room temperature. The suspension was recovered and the supernatant is removed. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 176nm with a  $d_{v10}$  and  $d_{v90}$  of 113nm and 248nm respectively without any agglomerate above 380nm. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is in the range 1.1 to 1.6. According to SEM observations 96% of individual particles present a cube-like shape. The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.006$  (see Table 3). The density of this powder has increased compared to the Example 1a powder: it is  $5.67 (\pm 0.05) \text{ g.cm}^{-3}$  and its specific surface area is  $10.5 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti ratio equals 1.00. The powder only contains 106ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.72% in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

#### Example 7 – Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out by mixing 2g of barium titanate produced according to procedure described in Example 1a with 30g of a  $\text{BaCl}_2$  and NaOH solution at pH=14. This solution was initially prepared by dissolving 13.6g of  $\text{BaCl}_2$  and 66g of NaOH in 1 litre of water. The suspension was placed in a 40ml hydrothermal autoclave, itself placed in an oven heated at 250°C. Reaching a temperature of 250°C inside the autoclave required 1 hour, the autoclave was then allowed to stay at this temperature for 2 hours and then to cool naturally to room temperature. The suspension was recovered and the supernatant was removed. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 155nm with a  $d_{v10}$  and  $d_{v90}$  of 117nm and 205nm respectively. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is in the range 1.1 to 1.3. According to SEM observations 91% of individual particles present a cube-like shape. The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.002$  (see Table 3). The density of this powder has increased compared to the Example 1a powder: it is already  $5.65 (\pm 0.01) \text{ g.cm}^{-3}$  after only 2 hours of post-treatment at 250°C and its specific surface area is  $11.3 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti ratio equals 1.00. The powder only contains 228ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 1.00 % attributed in the range 200 to 500°C to the presence of residual hydroxyls in the barium titanate structure.

#### Example 8 – Hydrothermal Post-treatment

The hydrothermal post-treatment was carried out using the same procedure as Example 7 except the hydrothermal autoclave was maintained at 250°C for 6 hours instead of 2 hours. The suspension was recovered and the supernatant was removed. The powder was washed with several liquids and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 214nm with a  $d_{v10}$  and  $d_{v90}$  of 147nm and 297nm respectively. Image analysis has been carried out in order to define number and volume particle size distributions (see Tables 4 and 5). Table 6 shows that the ratio of  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  determined by sedimentation and image analysis is in the range 1.1 to 1.3. According to SEM observations 87% of individual particles present a cube-like shape. The powder is of tetragonal crystal structure as an X-Ray diffraction peak splitting has been identified and  $c/a = 1.005$  (see Table 3). The density of this powder

has increased compared to the Example 1a powder, it is  $5.668 (\pm 0.005) \text{ g.cm}^{-3}$  and its specific surface area is  $10.2 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti equals 1.00. The powder only contains 104ppm of sodium compared to 500ppm of Example 1a. Finally, thermogravimetric analysis shows a weight loss of 0.67 % in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

#### Comparative Examples 9 and 10

##### Comparative Example 9 – Production of Powder in One Step

Instead of being fed to the SFTR, 150ml of gel prepared according to the procedure of Example 1a was placed in a 250ml hydrothermal autoclave, itself placed in an oven heated at 250°C. Reaching a temperature of 250°C inside the autoclave required 1 hour, the autoclave was then allowed to stay at this temperature for 19 hours and then to cool naturally to room temperature. The suspension was recovered and the supernatant was removed. The powder was washed in several steps and then dried. The powder was then characterized.

As presented in Table 1, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 594nm with a  $d_{v10}$  and  $d_{v90}$  of 248nm and 919nm respectively. Figure 11 is a SEM micrograph showing particles with a predominantly polyhedral shape and a very wide primary particle size distribution. The c/a ratio could not be calculated as the XRD peak was too difficult to fit accurately due to the wide particle size distribution. The density is  $5.871 (\pm 0.002) \text{ g.cm}^{-3}$  and the specific surface area is  $4.7 \text{ m}^2/\text{g}$  (see Table 2). The Ba/Ti ratio equals 1.00. The powder contains 71ppm of sodium after 19 hours but when the treatment is only performed for 6 hours 2000ppm of sodium is detected. Finally, thermogravimetric analysis shows a weight loss of 0.10 % in the range 200 to 500°C attributed to the presence of hydroxyls in the barium titanate structure. Despite its high density, this powder shows a much wider particle size distribution than the powders produced using our 2-step process as well as 6% of particles larger than  $1 \mu\text{m}$ .

##### Comparative Example 10 – One-Step Followed by Post Heat Treatment

Two grams of barium titanate produced according to the procedure described in Example 1a were treated for 1 hour in a pre-heated furnace at 600°C. The powder was then characterized.

As presented in Table 1 and illustrated by Figure 12, the volume particle size distribution determined by sedimentation shows  $d_{v50}$  of 112nm with a  $d_{v10}$  and  $d_{v90}$  of 57nm and 1068nm respectively with 13 % of particles larger than  $1 \mu\text{m}$ . The density is only  $5.22 (\pm 0.01) \text{ g.cm}^{-3}$  and the specific surface area is  $23 \text{ m}^2/\text{g}$  (see Table 2). Thermogravimetric analysis shows a weight loss of 0.45% in the range 200 to 500°C attributed to the presence of residual hydroxyls in the barium titanate structure.

Despite the low weight loss attributed to the residual hydroxyls, this particle size distribution of this powder shows 13% of particles larger than 1 $\mu$ m.

#### Example 11– Ceramic Body

5 The barium titanate powder prepared according to the method of Example 1a followed by Example 4 was used to prepare a ceramic body. The powder was used to make cylindrical rods 8-9.5mm in diameter by isostatic pressing at pressures between 1.15 - 108 and 1.5 -108Pa. These rods were sintered at different temperatures for a fixed time. Discs approximately 1mm thick were cut from the rods, flat parallel faces were ground with emery and gold electrodes were evaporated onto them. The  
10 resulting capacitor has satisfactory properties.

#### Example 12 – Production of MLCC

The same barium titanate powder was mixed with dopants and then with an aqueous solution of polyvinyl alcohol binder to prepare a slurry. Mixing and milling were carried out in water, using a ball mill with 2mm diameter stabilized zirconia balls.  
15 Ceramic green foils were tape-casted. After sieve printing of Ni electrodes the foils were stacked, laminated and cut to form MLCCs of the size 1206 with 150 active dielectric layers. The MLCCs were fired at 1300°C for 2h in a reducing atmosphere of moist nitrogen and hydrogen. The fired chips were terminated with Cu electrodes and then electroplated with nickel and tin to provide surface mountable MLCCs. The  
20 resulting capacitors have satisfactory properties.

#### Tables

Various characteristics of the Examples are shown in Tables 1 to 7. Table 8 compares characteristics of powders prepared according to the inventive process with those of EP 1 205 440.

TABLE 1

	<b>d<sub>v10</sub></b> <b>(nm)</b>	<b>d<sub>v50</sub></b> <b>(nm)</b>	<b>d<sub>v84</sub></b> <b>(nm)</b>	<b>d<sub>v100</sub></b> <b>(nm)</b>	<b>d<sub>v100</sub></b> <b>(nm)</b>	<b>Span</b>	<b><math>\sigma_g</math></b>	<b>% &gt;1<math>\mu</math>m</b>
Example 1a	48	75	-	116	-	0.91	-	0
Example 1b	101	155	-	252	-	0.97	-	0
Example 2	71	106	139	151	250	0.75	0.76	0
Example 3	85	130	172	189	365	0.80	0.76	0
Example 4	152	187	219	229	265	0.40	0.85	0
Example 5	105	154	200	217	350	0.73	0.77	0
Example 6	113	176	230	248	380	0.77	0.77	0
Example 7	117	155	-	205	-	0.57	-	0
Example 8	147	214	-	297	-	0.70	-	0
Example 9	248	594	-	919	-	1.13	-	6
Example 10	57	112	-	1068	-	9.00	-	13

Table 1 presents the characteristics of the volume particle size distributions determined by sedimentation. Examples 2 to 8 prepared according to the described novel method show median volume diameter  $d_{v50}$  in the range 0.106 to 0.214 $\mu$ m, a span in the range 0.40 to 0.91, a geometric standard deviation  $\sigma_g$  in the range 0.76 to 0.85 and no agglomerates above 1 $\mu$ m.

Table 2 presents the density and specific surface area, and the BET diameter  $d_{BET}$  and the factor of agglomeration  $F_{AG}$  calculated using those data are also presented. Examples 2 to 8 prepared according to the described novel method show a density in the range 5.58 to 5.73 g.cm<sup>-3</sup> corresponding to 92.8 to 95.3% of the intrinsic density of large crystals of barium titanate. Also for those Examples, the  $F_{AG}$  is in the range 1.6 to 2.1.

TABLE 2

Examples	Density (g.cm <sup>-3</sup> )	Specific surface area (m <sup>2</sup> .g <sup>-1</sup> )	d <sub>BET</sub> (nm)	F <sub>AG</sub>
Example 1a	5.0	-	-	-
Example 1b	5.0	-	-	-
Example 2	5.60 (±0.09)	-	-	-
Example 3	5.58 (±0.10)	-	-	-
Example 4	5.73 (±0.06)	11.6	90	2.1
Example 5	5.71 (±0.01)	12.1	87	1.8
Example 6	5.67 (±0.05)	10.5	101	1.7
Example 7	5.65 (±0.01)	11.3	94	1.6
Example 8	5.668 (±0.005)	10.2	104	2.1
Example 9	5.871 (±0.002)	4.7	217	2.7
Example 10	5.22 (±0.01)	23	50	2.2

Table 3 presents the value of weight losses attributed to removal of hydroxyls as measured by thermogravimetric (TGA) analysis, the c/a ratios and the sodium concentrations. The weight losses attributed to removal of hydroxyls vary between 0.67 and 1.0 % depending upon the parameters of the hydrothermal post-treatment. The c/a ratios of Examples 2 to 8 all correspond to a tetragonal structure.

Table 4 exhibits the number particle size distribution defined from image analysis of SEM micrographs. The median number diameter  $d_{n50}$  for Examples 2 to 8 is in the range 0.082 to 0.105  $\mu\text{m}$ , while the corresponding span is in the range 0.29 to 0.43 and the geometric standard deviation  $\sigma_g$  is in the range 0.84 to 0.92.

TABLE 3

Examples	TGA (200 to 500 °C) Wt %	XRD c/a	Na (ppm)
Example 1a	4.2	1.00	500
Example 1b	-	1.00	-
Example 2	0.8	1.002	34
Example 3	0.7	1.006	26
Example 4	0.7	1.007	31
Example 5	0.8	1.006	65
Example 6	0.72	1.006	106
Example 7	1	1.002	228
Example 8	0.67	1.005	104
Example 9	0.1	•	71
Example 10	0.45	1.00	~500

- The c/a ratio was not calculated: the XRD peak was too difficult to fit accurately due to the wide particle size distribution.

TABLE 4

	d <sub>n10</sub> nm	d <sub>n25</sub> nm	d <sub>n50</sub> nm	d <sub>n75</sub> nm	d <sub>n84</sub> nm	d <sub>n90</sub> nm	d <sub>n100</sub> nm	Span	$\sigma_g$	QR
Example 2	68	75	82	91	96	99	133	0.38	0.85	1.21
Example 3	76	82	92	102	106	112	200	0.39	0.87	1.24
Example 4	80	86	96	108	114	118	144	0.40	0.84	1.26
Example 5	74	82	94	103	106	110	122	0.38	0.89	1.26
Example 6	76	86	98	106	111	114	167	0.39	0.88	1.23
Example 7	82	91	98	105	107	110	144	0.29	0.92	1.15
Example 8	83	92	105	115	122	128	178	0.43	0.86	1.25
Example 9	140	176	238	331	390	440	1000	1.26	0.61	1.88

TABLE 5

	<b>d<sub>v10</sub></b> <b>(nm)</b>	<b>d<sub>v25</sub></b> <b>(nm)</b>	<b>d<sub>v50</sub></b> <b>(nm)</b>	<b>d<sub>v75</sub></b> <b>(nm)</b>	<b>d<sub>v84</sub></b> <b>(nm)</b>	<b>d<sub>v90</sub></b> <b>(nm)</b>	<b>d<sub>v100</sub></b> <b>(nm)</b>	Span	$\sigma_g$	QR
Example 2	90	99	109	120	125	132	165	0.39	0.87	1.21
Example 3	101	111	124	142	160	178	248	0.62	0.78	1.28
Example 4	104	115	130	142	148	154	208	0.38	0.88	1.23
Example 5	100	112	124	133	136	140	156	0.32	0.91	1.19
Example 6	104	117	129	140	146	154	208	0.39	0.88	1.20
Example 7	104	117	129	140	146	154	194	0.39	0.88	1.20
Example 8	112	126	140	157	168	180	220	0.49	0.83	1.25
Example 9	254	340	506	698	848	978	1000	1.43	0.60	2.05

TABLE 6

<b>Examples</b>	<b>d<sub>v10</sub></b> <b>ratio</b>	<b>d<sub>v50</sub></b> <b>ratio</b>	<b>d<sub>v90</sub></b> <b>ratio</b>
Example 2	0.8	1.0	1.1
Example 3	0.8	1.0	1.1
Example 4	1.5	1.4	1.5
Example 5	1.1	1.2	1.6
Example 6	1.1	1.4	1.6
Example 7	1.1	1.2	1.3
Example 8	1.3	1.5	1.7
Example 9	1.0	1.2	0.9

TABLE 7

<b>Examples</b>	<b>Shape of primary particles</b>	<b>Structure</b>	<b>Span (sedimentation)</b>
Example 1a	sphere	Cubic	0.91
Example 1b	sphere	Cubic	0.97
Example 2	85% cube	Tetragonal	0.75
Example 3	85% cube	Tetragonal	0.80
Example 4	90% cube	Tetragonal	0.40
Example 5	86% cube	Tetragonal	0.73
Example 6	96% cube	Tetragonal	0.77
Example 7	91% cube	Tetragonal	0.57
Example 8	87% cube	Tetragonal	0.70
Example 9	polyhedra	Undetermined	1.13
Example 10	generally sphere	Cubic	9.00

Table 5 exhibits the volume particle distribution defined from image analysis of SEM micrographs. The median volume diameter  $d_{v50}$  for examples 2 to 8 is in the range 0.109 to 0.140 $\mu$ m.

- 5 Table 6 compares the  $d_{v10}$ ,  $d_{v50}$  and  $d_{v90}$  ratios by sedimentation/image analysis. As shown in Table 6, the ratios between the sedimentation  $d_{v50}$  and the image analysis  $d_{v50}$  is in the range 1.0 to 1.5, confirming the absence of agglomerates made up of numerous (more than 8) primary particles.

- 10 Table 7 presents the shape observed for each example, including the amount in % of particles of cube-like shape encountered for Examples 2 to 8. All those examples also show a tetragonal structure and a span of the particle size distributions (determined by sedimentation) smaller than 1.

- 15 Table 8 compares various characteristics of powders prepared according to the method of this invention with those of EP 1 205 440, highlighting notably the different particle shapes, and the conversion from cubic structure obtained after a rapid low-temperature first stage reaction to tetragonal by the solvo- or hydrothermal second stage below 400°C according to the invention, compared to conversion from cubic structure obtained after ageing and hydrothermal treatment in EP 1 205 440 followed by calcination at 500-1000°C to obtain tetragonal structure.

Table 8:

Characteristics	EP 1,205,440	This invention
Average size or $d_{n50}$ ( $\mu\text{m}$ ) by image analysis	0.05-0.5 $\mu\text{m}$	0.005 to 0.25 $\mu\text{m}$
Geometric standard deviation ( $\sigma_g = d_{n50}/d_{n84}$ ) of particle size distribution determined by image analysis	$\sigma_g \geq 0.70$	$0.75 \leq \sigma_g < 1$
Particle shape	Spherical with sphericity (maximum diam/minimum diameter) from 1.0 to 1.4	Isotropic and faceted individual particle shape, in particular a cube-like shape
Crystal structure	Cubic structure after ageing and hydrothermal treatment, converted to tetragonal by calcination at 500-1000°C	Cubic structure after rapid low temperature first stage, converted to tetragonal by a second stage below 400°C
c/a	Without coating: $1.005 < c/a < 1.014$ With coating: $1.007 < c/a < 1.016$	$1.002 < c/a < 1.008$
Specific Surface Area ( $\text{m}^2/\text{g}$ )	2-20	2-30
Density	High denseness after calcining	High density after solvo- or hydrothermal treatment, corresponding to at least 90% of the intrinsic density of a large crystal

All of the Examples above are described on laboratory scale. To scale up for industrial production, the powder of the first step of the process can be prepared in an SFTR apparatus with multiple reaction tubes, however while maintaining suitable small segmented reaction volumes. On the other hand, the hydrothermal post  
5 treatment can be scaled up by increasing the size of the reactor/autoclave or by using a continuous hydro- or solvothermal apparatus. This opens new possibilities for producing ultrafine barium titanate on a large scale, by producing the ultrafine particles of defined characteristics but low density from small reaction volumes in the first stage, then converting large amounts of the low density ultrafine particles into  
10 high density ultrafine particles during the second stage.